



# Cambridge IGCSE™

**CHEMISTRY**

**0620/12**

Paper 1 Multiple Choice (Core)

**February/March 2022**

**45 minutes**

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

## INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

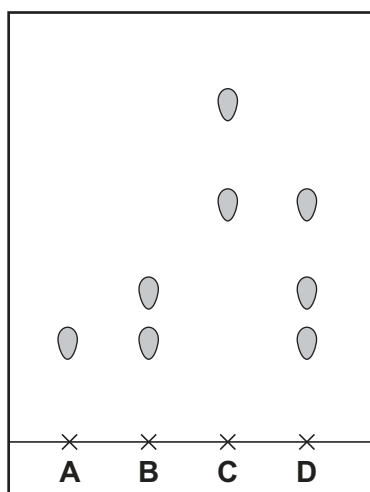
## INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages. Any blank pages are indicated.



- 1 Which change of state is an exothermic process?
- A condensation
  - B evaporation
  - C melting
  - D sublimation
- 2 In which state does  $1 \text{ dm}^3$  of methane contain the most particles?
- A gas at  $100^\circ\text{C}$
  - B gas at room temperature
  - C liquid
  - D solid
- 3 Which dye on the chromatogram is a pure substance?



- 4 Which piece of apparatus is used to measure exactly  $5.00 \text{ cm}^3$  of a liquid?
- A  $5 \text{ cm}^3$  beaker
  - B  $10 \text{ cm}^3$  measuring cylinder
  - C  $25 \text{ cm}^3$  pipette
  - D  $50 \text{ cm}^3$  burette

- 5 Fermentation of sugar produces a mixture of ethanol solution and solid yeast.

How is the solid yeast removed from the mixture?

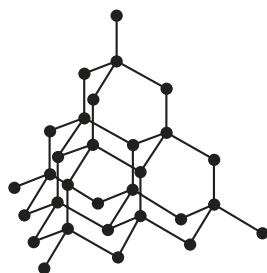
- A crystallisation
- B distillation
- C filtration
- D fractional distillation

- 6 Matter exists as elements, compounds and mixtures.

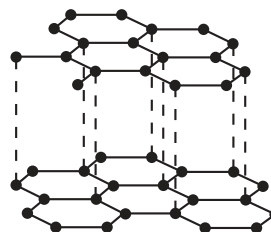
Which row identifies an element, a compound and a mixture?

	element	compound	mixture
<b>A</b>	calcium	potassium carbonate	sodium chloride
<b>B</b>	brass	sodium chloride	air
<b>C</b>	calcium	sodium chloride	brass
<b>D</b>	sodium chloride	water	potassium carbonate

- 7 Which pair of statements about diamond and graphite is correct?



diamond



graphite

- A Diamond and graphite are both pure carbon. They are both macromolecules.
- B Diamond and graphite can both be used as electrodes. Graphite is also used as a lubricant.
- C Diamond has covalent bonds. Graphite has ionic bonds.
- D Diamond is hard with a high melting point. Graphite is soft with a low melting point.

- 8 An isotope of chromium is represented by  ${}_{24}^{52}\text{Cr}$ .

Which statement about an atom of this isotope of chromium is correct?

- A** It contains 24 electrons.  
**B** It contains 24 neutrons.  
**C** It contains 28 protons.  
**D** It contains 52 neutrons.

- 9 Sodium is in Group I of the Periodic Table and chlorine is in Group VII.

Which row describes what happens when sodium bonds ionically with chlorine?

	sodium atoms	ion formed	chlorine atoms	ion formed
<b>A</b>	gain an electron	$\text{Na}^-$	lose an electron	$\text{Cl}^+$
<b>B</b>	gain an electron	$\text{Na}^+$	lose an electron	$\text{Cl}^-$
<b>C</b>	lose an electron	$\text{Na}^-$	gain an electron	$\text{Cl}^+$
<b>D</b>	lose an electron	$\text{Na}^+$	gain an electron	$\text{Cl}^-$

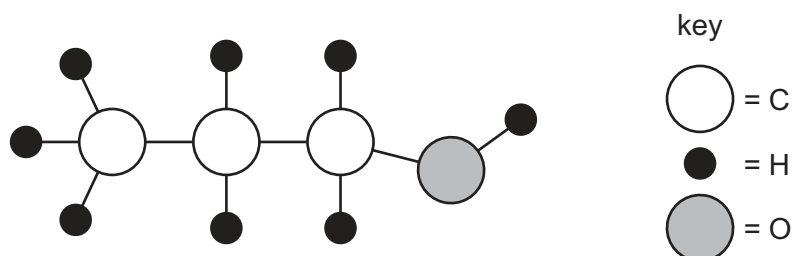
- 10 Caesium fluoride is an ionic compound.

Which statements about caesium fluoride are correct?

- It conducts electricity when solid.
- It has a high melting point.
- It is soluble in water.
- It is highly volatile.

- A** 1 and 2      **B** 1 and 4      **C** 2 and 3      **D** 3 and 4

- 11 The structure of a molecule of a compound is shown.



What is the formula of this compound?

- A**  $\text{C}_3\text{H}_7\text{O}$       **B**  $\text{C}_3\text{H}_8\text{O}$       **C**  $\text{C}_8\text{H}_3\text{O}$       **D**  $\text{C}_8\text{HO}_3$

- 12 Calcium carbonate,  $\text{CaCO}_3$ , reacts with dilute hydrochloric acid to produce carbon dioxide.

The equation for the reaction is shown. The relative formula mass of calcium carbonate is 100.



10g of calcium carbonate is reacted with an excess of dilute hydrochloric acid.

Which mass of carbon dioxide is produced?

- A** 2.2g      **B** 2.8g      **C** 4.4g      **D** 44g

- 13 Molten sodium chloride and concentrated aqueous sodium chloride are electrolysed using platinum electrodes.

What are the products at the negative electrode (cathode) in each electrolysis?

	molten sodium chloride	concentrated aqueous sodium chloride
<b>A</b>	hydrogen	hydrogen
<b>B</b>	hydrogen	sodium
<b>C</b>	sodium	hydrogen
<b>D</b>	sodium	sodium

- 14 An object is electroplated with silver using an aqueous silver salt as the electrolyte.

Which row is correct?

	the object to be electroplated is the	the other electrode is made from
<b>A</b>	anode	carbon
<b>B</b>	anode	silver
<b>C</b>	cathode	carbon
<b>D</b>	cathode	silver

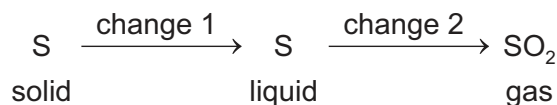
15 Which row describes the changes that occur in an endothermic reaction?

	energy change	temperature
<b>A</b>	energy given out to the surroundings	decreases
<b>B</b>	energy given out to the surroundings	increases
<b>C</b>	energy taken in from the surroundings	decreases
<b>D</b>	energy taken in from the surroundings	increases

16 Which statement about fuels is correct?

- A** Heat energy is only produced by burning fuels.
- B** Hydrogen is used as a fuel although it is difficult to store.
- C** Methane is a good fuel because it produces only water when burned.
- D** Uranium is burned in air to produce energy.

17 A sequence of changes involving sulfur is shown.



Which row describes the changes?

	change 1	change 2
<b>A</b>	chemical	chemical
<b>B</b>	chemical	physical
<b>C</b>	physical	chemical
<b>D</b>	physical	physical

18 Magnesium is added to dilute hydrochloric acid.

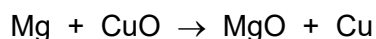
25 cm<sup>3</sup> of gas is given off in the first 30 s of the reaction.

The experiment is repeated at a lower temperature. All other reaction conditions are the same.

Which volume of gas is produced in the first 30 s of this reaction?

- A** 15 cm<sup>3</sup>
- B** 25 cm<sup>3</sup>
- C** 30 cm<sup>3</sup>
- D** 50 cm<sup>3</sup>

19 The equation for the reaction between magnesium and copper(II) oxide is shown.



Which substance is oxidised?

- A** Cu                      **B** CuO                      **C** Mg                      **D** MgO

20 Methyl orange is added to dilute hydrochloric acid and to aqueous sodium hydroxide.

What is the colour of the methyl orange in each solution?

	colour in dilute hydrochloric acid	colour in aqueous sodium hydroxide
<b>A</b>	orange	red
<b>B</b>	red	yellow
<b>C</b>	red	orange
<b>D</b>	yellow	red

21 Compound X is dissolved in water and two separate samples of the solution are tested.

The results of the tests are shown.

test	observation
add aqueous sodium hydroxide	a white precipitate forms which is insoluble in excess
acidify with dilute nitric acid and add aqueous silver nitrate	a yellow precipitate forms

What is compound X?

- A** calcium chloride  
**B** calcium iodide  
**C** zinc chloride  
**D** zinc iodide

22 Which statement about the Periodic Table is correct?

- A** Elements with the highest atomic number in each period are metallic.  
**B** Elements with the lowest group numbers are non-metals.  
**C** Elements with similar chemical properties are placed in groups.  
**D** Elements with similar physical properties are placed in periods.





26 Which process is used to obtain the metal calcium from its ore?

- A electrolysis
- B oxidation with carbon
- C reduction with carbon
- D thermal decomposition

27 Which row links the property of a metal to its use?

	property	use
A	high density	aircraft bodies
B	high reactivity	food containers
C	good electrical conductor	cooking pans
D	ductile	electrical wiring

28 The table gives some properties of an element.

melting point in °C	3422
appearance of the element	grey
appearance of the chloride of the element	dark blue
density in g / cm <sup>3</sup>	19.2
electrical conductivity when solid	good

Which other property does this element have?

- A acts as a catalyst
- B brittle
- C forms an acidic oxide
- D highly reactive with water

29 A metal reacts vigorously with cold water.

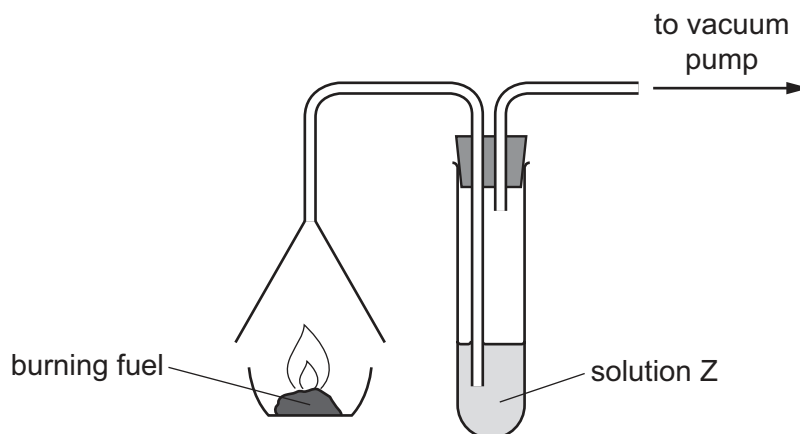
Which statement about the metal is correct?

- A It is above hydrogen in the reactivity series.
- B It is below magnesium in the reactivity series.
- C Its oxide can be reduced with carbon.
- D It does not react with dilute acids.

- 30 Which row describes the colour changes when water is added to anhydrous cobalt(II) chloride and anhydrous copper(II) sulfate?

	anhydrous cobalt(II) chloride	anhydrous copper(II) sulfate
<b>A</b>	blue to pink	white to blue
<b>B</b>	blue to white	blue to pink
<b>C</b>	pink to blue	blue to white
<b>D</b>	white to blue	pink to blue

- 31 The gases produced by a burning fuel are passed through solution Z using the apparatus shown. The fuel contains compounds of sulfur.



Which row identifies solution Z and the result obtained when the fuel contains compounds of sulfur?

	solution Z	result
<b>A</b>	acidified potassium manganate(VII)	turns colourless
<b>B</b>	acidified potassium manganate(VII)	turns purple
<b>C</b>	litmus	bleached
<b>D</b>	litmus	turns blue

32 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane	
<b>A</b>	formed when vegetation decomposes	✓	✗	key ✓ = correct ✗ = not correct
<b>B</b>	greenhouse gas	✓	✓	
<b>C</b>	present in unpolluted air	✗	✗	
<b>D</b>	produced during respiration	✗	✓	

33 Which row identifies uses of sulfur?

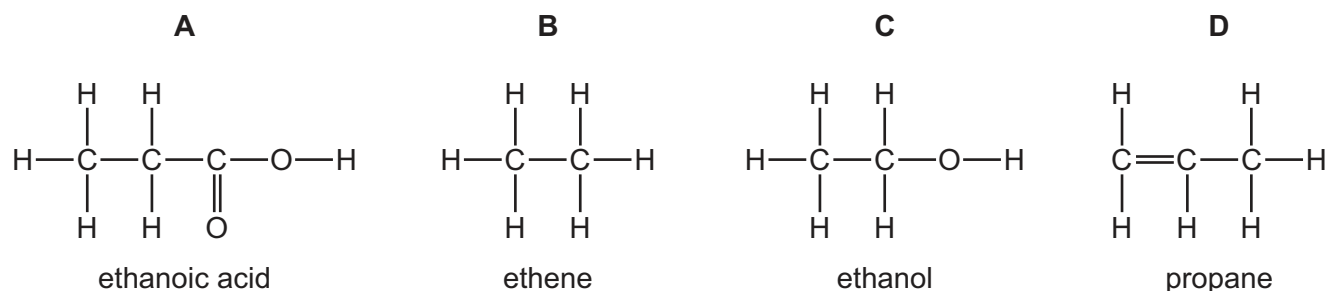
	use 1	use 2	use 3
<b>A</b>	making ammonia	bleaching wood pulp	food preservative
<b>B</b>	making sulfuric acid	bleaching wood pulp	food preservative
<b>C</b>	making sulfuric acid	food preservative	as an NPK fertiliser
<b>D</b>	making ammonia	food preservative	as an NPK fertiliser

34 Which statements about lime are correct?

- 1 Lime is made by heating calcium carbonate (limestone).
- 2 Lime is used to desulfurise flue gases.
- 3 Lime is used to treat alkaline soil.
- 4 The chemical name for lime is calcium oxide.

**A** 1 and 3      **B** 1, 2 and 4      **C** 1 and 4 only      **D** 2, 3 and 4

35 Which structure is correctly named?



36 The fractional distillation of petroleum produces a series of fractions with different uses.

Which row identifies a use for a fraction?

	fraction	use
<b>A</b>	bitumen	jet fuel
<b>B</b>	gas oil	cooking
<b>C</b>	kerosene	making roads
<b>D</b>	naphtha	making chemicals

37 Ethene and propene are both members of the same homologous series.

Which statements explain why ethene and propene have similar chemical properties?

- 1 They are both hydrocarbons.
- 2 They are both made by cracking.
- 3 They have the same functional group.

**A** 1 and 2      **B** 1 and 3      **C** 2 only      **D** 3 only

38 Which statement about ethane is correct?

- A** It decolourises bromine water.
- B** It burns in excess oxygen to form water and carbon dioxide.
- C** Its molecular formula is  $C_2H_4$ .
- D** Its atoms are joined together by ionic bonding.

39 Which statements about ethanol are correct?

- 1 Ethanol is used as a solvent.
- 2 Ethanol can be made directly from ethane.
- 3 Ethanol is a covalent compound.

**A** 1 only      **B** 1 and 2      **C** 1 and 3      **D** 2 and 3

**40** Polymers are long-chain molecules made from small molecules linked together.

Four polymers or types of polymer are listed.

- 1 carbohydrates
- 2 nylon
- 3 proteins
- 4 *Terylene*

Which polymers or types of polymer are synthetic?

- A** 1 and 3      **B** 1 and 4      **C** 2 and 3      **D** 2 and 4



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The Periodic Table of Elements

		Group															
I	II	III	IV	V	VI	VII	VIII					VIII					
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<div style="border: 1px solid black; padding: 5px; text-align: center;"> <b>Key</b>                      atomic number                      atomic symbol                      name                      relative atomic mass                 </div>										2 <b>He</b> helium 4					
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24											5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	—	—	—	—

lanthanoids	57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
actinoids	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).